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A greener approach to synthesizing metal-decorated carbogels from alginate for emerging technologies

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Abstract

In the present work, a series of metal nanoparticle-decorated carbogels (M-DCs) was synthesized starting from beads of parent metal-crosslinked alginate aerogels (M-CAs), comprising Ca(II), Ni(II), Cu(II), Pd(II) and Pt(IV) ions, followed by pyrolysis under N_2 atmosphere up to pyrolysis temperatures of T_P = 600 ºC. The textural properties of M-CAs are found to depend on the crosslinking ion, yielding fibrous pore networks with high specific mesoporous volume and specific surface area S_V (S_V ~ 480 – 687 m^2/g) for M-CAs crosslinked with hard cations, Ca(II), Ni(II) and Cu(II), and comparably loose networks with increased macroporosity and lower specific surface (S_V ~ 240 – 270 m^2/g) for Pd(II) and Pt(IV) crosslinked aerogels. The pyrolysis of M-CAs resulted in two simultaneously occurring processes: changes of the solid backbone and the growth of metal/metal oxide nanoparticles (NPs). The thermogravimetric analysis (TGA) showed a significant influence of the crosslinking cation on the decomposition mechanism and associated change of textural properties. Backscattered electron scanning electron microscopy imaging (SEM-BSE) and X-ray diffraction revealed that metal ions (molecularly dispersed in the parent aerogels) formed nanoparticles composed of elementary metals and metal oxides in varying ratios over the course of pyrolytic treatment. Increasing the T_P led to in generally larger nanoparticles. The pyrolysis of the nickel-crosslinked aerogel (Ni-CA) preserved to a large extent the mesoporous structure and resulted in the evolution of fine (~ 14 nm) homogeneously dispersed Ni/NiO nanoparticles. Overall, this work presents a green approach for synthesizing metal-nanoparticle containing carbon materials, useful in emerging technologies related to heterogeneous catalysis, electrocatalysis among others.

Keywords: *green chemistry, aerogels, carbogels, alginate crosslinking, pyrolysis treatment, metallic-clusters evolution*

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Abbreviations

M-CA: metal-crosslinked aerogel

M-DC: metal nanoparticle-decorated carbogels

M: Ca, Ni, Cu, Pd or Pt

NPs: Nanoparticles

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Introduction

Since the discovery of aerogels in the 1930s, a special attention has been paid to their development for a number of industrial applications.[1,2] Given their nanoporous, thermal and electrical properties, breakthroughs in the past decades have been made in the sector of foods, biomedical, wastewater treatment, pollution, construction, electrode materials and catalysis.[3,4] Since the late 1980s carbon aerogels (carbogels) have presented a promising alternative for natural carbonaceous materials, due to the possibility of controlling porosity, purity and mechanical strength by adjustments of the synthesis pathway.[5] The tailorable textural and structural features of carbogels allow us to overcome some constraints related to metal-nanoparticles materials, such as formation of metal-species aggregates, pore blocking and leaching. Other features such as large specific area, low density, low thermal and high electrical conductivity as well as hydrophobicity make them attractive materials as catalyst supports, chemical adsorbents and energy storage as capacitors.[6] The polycondensation of resorcinol with formaldehyde under alkaline conditions was a typical pathway to obtain carbogels in recent decades. Other synthetic polymers, such as polybenzoxazines, polyureas, polyimides, polyamides and polyurethanes, have also been used as porous precursors for carbogels.[7] For instance, Raptopoulos *et al*. produced carbogels from polyurea crosslinked alginate, with an increased carbonization yield of up to 37%, which in turn improves surface functionalities (*e.g.* electrical conductivity and catalytic activity) of the carbogel by contributing with oxygen and nitrogen functional groups.[8] Recently, Zhang *et al*. developed a series of carbon aerogel-based electrocatalysts with transition metal-nitrogen sites for the electroreduction of CO_2.[9] The starting hydrogel was formed *via* complexation between 2,2′-bipyridine-4,4′-dicarboxylic acid and selected metal ions (Ni, Fe, Co, Mn and Cu) followed by physical crosslinking by polypyrrole. Baumann *et al*. incorporated copper in a carbogel matrix from formaldehyde polymerization, achieving a uniform dispersion of spherical metal nanoparticles, ranging from 10 to 50 nm.[10] In another study, potassium-doped hydrogels were synthesized via formaldehyde polymerization with the potassium salt of 2,4-dihydroxybenzoic acid: These hydrogels were formed on monodisperse polystyrene particle

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templates (diameter = 300 or 450 nm). Subsequent ion exchange with Ni, Co or Cu-salts and supercritical drying followed by pyrolysis at 800 °C then led to macroporous ordered carbons.[11] The materials obtained presented dense networks of interconnected carbon particles with hierarchical pore structure of meso- and microporosity in the wall that facilitates the accessibility to the metal sites. In the conventional synthesis of metal-containing carbonaceous 3D structures, the incorporation of metals requires subsequent post-functionalization, leading to the need for additional process steps and resource consumption. As an example, a sulphonated starch-based support, is first pyrolysed at 400 ºC and subsequently impregnated with palladium acetate (10% by weight) in acetone.[12] Wet impregnation, along with other elaborate methods, e.g., chemical vapor deposition[13] represent additional challenges such as: i) post-distributing metal clusters homogeneously through the adhesive-resistant carbon network, and ii) aggregation of dispersed metals on the carbon surface resulting in the formation of large metal cluster. Despite the remarkable properties exhibited by materials generated via the aforementioned methods, they rely on expensive chemicals and have adverse environmental impacts. Consequently, there has been a growing interest in exploring alternative carbon sources, particularly non-expensive and abundant polysaccharides. Especially polysaccharides derived from species with a large capability to fix carbon dioxide, such as the brown algae, have gained much interest as precursors for the synthesis of nanoporous carbon materials.[14,15] From a sustainable perspective, this CO_2 capture feature can add value to novel algae-derived nanomaterials. On average, brown algae species contains ~ 40 % alginate, an anionic biopolymer that can be easily crosslinked with cations for the preparation of hydrogels.[16,17] For example, Meng *et al.* crosslinked sodium alginate with $CaCl_2$ to obtain a gel that was dried at 50 ºC for 3 h, and pyrolized at temperatures up to 1000 ºC, under nitrogen atmosphere.[18] The resulting carbon aerogel was soaked in 2M of HCl and washed with deionized water to obtain a near neutral sample. In the synthesis of cellulose-based carbogels for the removal of heavy metals from wastewater, Alatalo *et al*. employed concentrated sulfuric acid to catalyze the degradation of cellulose prior to hydrothermal carbonization.[19]

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The above reports evidence the need for consolidating the production of carbogels based on inexpensive and more abundant biomass-derived carbon sources, while minimizing the use of no-green chemicals during the process. On the basis of the principles of green chemistry and modern aerogel science, the present work aims at developing an easy, flexible, and inexpensive strategy towards green and sustainable metal-decorated carbogels. The selection of alginate as the carbon source allowed the reduction of production costs and carbon footprint. Therefore, a series of metal nanoparticle-decorated carbogels (M-DCs) were synthesized, starting from the beads of parent metal-crosslinked aerogels (M-CAs), comprising the ions Ni(II), Cu(II), Pd(II), Pt(IV). Because Ca(II) is the most commonly used ion for alginate crosslinking and in the alginate aerogel production, Ca-crosslinked gels were included in the study as a reference material. The *in situ* generation of monodispersed metal-nanoparticles was controlled by the pyrolysis temperature, with no additional post-modification step. The conversion of the biopolymer-aerogel network into the corresponding carbogels was studied with a focus on the changes in textural characteristics during the transition and the evolution of the metallic clusters in the range of low pyrolysis temperatures < 600 °C.

Materials and methods

Chemicals

The fine powder of sodium alginate Hydagen 500, was used as received from BASF (Ludwigshafen, Germany). The metal precursors were: calcium chloride dihydrate ($CaCl_2 \cdot 2H_2O$) (>99%, Carl Roth catalog number: T885.3), nickel (II) nitrate hexahydrate (>99%, Bernd Kraft, catalog number: 15306), copper (II) sulfate pentahydrate (>99%, Riedel-de Haën), palladium (II) nitrate dihydrate (~40% Pd basis; Sigma Aldrich, catalog number: 76070-1G) and platinum (IV) chloride (assay 96%; Sigma Aldrich, catalog number: 206113-1G). All chemicals were used without further purification.

Synthesis of metal-crosslinked aerogels (M-CAs)

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A 2 wt% aqueous solution of alginate was prepared by adding sodium alginate powder to demineralized water. The solution was first stirred for 10 min at 500 rpm and then stirred at 1000 rpm up to 2 h at room temperature. Before further usage, the solution was stored at room temperature for 12 h to allow the solution to settle. Hydrogel particles were prepared by adding 35 g of alginate solution to a dripping device (**Fig. S1**) that consisted of a separating funnel connected to a compressed air supply and equipped with a nozzle made from a micropipette tip (10 mL) and a custom cut syringe needle (0.6 mm ID and 3 mm length). The falling distance (approx. 7 cm from the gelation bath) and compressed air pressure (< 0.5 bar) were kept constant throughout all experiments. Nickel, copper, palladium, platinum and calcium containing aerogels were prepared by dripping the alginate solution in an aqueous gelation bath with a corresponding salt, where droplets quickly formed hydrogels when coming into contact with a gently stirred gelation salt solution. Three different concentrations c_m of Ca(II), Ni(II) and Cu(II) metal salts were set in the gelation bath (c_m = 17, 180 and 380 mmol/L), Pd(II) and Pt (IV) salts were used at a constant c_m (17 mmol/L). After a gelation time of 20 min, the gels were separated from the gelation solution, washed twice with demineralized water and several times with ethanol (99.8%) until a final ethanol concentration ≥ 96 wt% was achieved. According to the previous works the Pd(II)-crosslinked gels showed a color change to orange-brown when soaked in ethanol. This was attributed to the reduction of a portion of Pd(II).[20] However, this phenomenon was proved to have no effect on the formation or an the final structure the gel. The alcogels were sealed in filter bags and transferred into a high-pressure vessel (3.9 L) for supercritical drying (70 ºC, 120 bar) under a constant CO_2 flow (120 g/min) for 6 h. The dried M-CA beads were collected after slow depressurization (1 bar/min) and stored in a desiccator for further analysis and pyrolysis.

Preparation of metal nanoparticle-decorated carbogels (M-DCs) and thermogravimetric analysis

Thermogravimetric analysis (TGA) of M-CAs and carbogel preparation were carried out simultaneously in the same TGA equipment (Linseis STA-PT1600) under a flow of 100 cm^3/min of

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nitrogen. The samples were heated up to 90 ºC at a heating rate of 5 °C/min, followed by a dwelling phase of 45 min. Subsequently, the temperature was raised to the desired final temperature at a heating rate of 5 °C/min, followed by an additional dwelling time of 60 min. Six different final pyrolysis temperatures T_P (150, 285, 380, 450, 525 and 600 °C) were set as part of the thermal stability analysis and M-DC production. For both Pd and Pt containing gels, temperatures of 265 °C and 365 ºC were applied instead of 285 °C and 380 ºC, respectively. Between 17 mg and 161 mg of each sample was used, depending on the T_P.

Characterization of M-CAs and M-CDs

The metal content was determined via ICP-OES using a Perkin Elmer Optima 8300 DV. Samples were previously dissolved in a combination of HNO_3, HCl and HF solution. The textural properties of M-CAs and M-DCs were analyzed by N_2 adsorption-desorption measurements (Nova 3000e Surface Area Analyzer Quantachrome Instruments, Boynton Beach, FL, USA). An overall sample mass of 20–30 mg was used and all samples were degassed under vacuum at 60 °C for at least 6 h prior to each analysis. The Brunauer-Emmett-Teller (BET) method was used to estimate the specific surface area (S_V), the pore volume of the mesopores (V_{meso}) and pore diameter (d_{meso}) were determined by the Barrett-Joyner-Halendia (BJH) method. All gas sorption experiments were carried out as single determinations with an estimated relative measurement error of 5%. The particle size and sphericity SPH of M-CA beads were determined using a Camsizer XT system (Retsch Technology) in free fall mode according to our previous works.[21] Skeletal densities ρ_s were determined via helium pycnometry (Multivolume Micromeritics 1305, 4-fold determination). The bulk densities were calculated from the weight of beads that filled a graduated cylinder (volume = 0.35 cm^3). The total porosity ε was estimated based on former work[22], as follows:

$$\varepsilon = (1 - \rho_b/\rho_s)\cdot 100\%$$

where ρ_b is the bulk density, and ρ_s is the skeletal density.

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The outerskin and inner pore structure morphologies of the M-CAs were characterized by scanning electron microscopy (SEM, Zeiss Supra VP55, Jena, Germany). The measurements were carried out under high vacuum at an accelerating voltage of 3 – 5 kV.

Reducibility and oxidizability of Pd-CA and M-DCs were determined by means of temperature programmed reduction (H_2-TPR). In a pretreatment step, the sample was heated at a rate of 10 °C/min to 150 °C, maintained for 0.5 h under a constant flow (50 cm^3/min) of pure argon and finally returned to room temperature to start the data collecting. The H_2-TPR experiments were conducted up to temperatures of 1000 ºC under a constant flow of H_2/Ar (10% v/v; 50 cm^3/min) at a heating rate of 10 ºC/min, in a Autochem II 2920 Micromeritcs equipment, coupled to a quadrupole mass spectrometer (Pfeiffer-Vacuum Modelo Omnistar). Hydrogen consumption was monitored by the thermal conductivity detector. Before the detector, an ice trap was used to retain any water formed during the reduction.

Characterization of M-DCs via backscattered electron imaging and X-ray diffraction

In order to image the metallic nanoparticles (NPs) in M-DCs obtained at different pyrolysis temperatures, low voltage scanning electron microscopy was performed using a ThermoFisher Scientific Scios 2 microscope. Backscattered electrons (BSE) were detected. BSE electrons are from a deeper interaction volume of the incident beam. The intensity of BSE shows a significant dependence on the atomic number, which results in a contrast between the elements in the BSE images. The contrast correlates with atomic number: brighter regions correspond to higher atomic numbers. The electron beam resolution was 1.6 nm at the optimum working distance < 5 mm at 2 kV accelerating voltage. The samples were fixed with a vacuum-resistant carbon tape on the sample holder. Sputter coating was not applied in this case because even a 5 nm thick sputtered gold layer can dramatically alter the morphology and the surface structure of the aerogel.[23]

X-ray diffraction patterns of M-DCs samples were collected under Cu-Kα1 radiation (Cu-Kα1 = 0.154059 nm) using a Rigaku SmartLab 9 kW X-ray diffractometer (XRD) with the following measurement conditions: measurements were carried out in Bragg-Brentano (Θ/2Θ) mode, between

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35 and 80° using a scan rate of 5°/min. The average crystallite size for each component of the NPs was calculated from the XRD data. Quantitative analysis of the XRD pattern was carried out using Rietveld refinement using the SmartStudio II software. In this process, the measured XRD profile is approximated by a calculated one that includes all instrumental and structural parameters. This method applies a non-linear least squares method based on a wide range of free parameters such as unit cell dimension, crystal structure, and peak shape. Using this fitted graph, quantitative information can be deduced from the XRD pattern.

RESULTS AND DISCUSSION

Properties of metal ion crosslinked aerogels (M-CAs)

The dripping method using bivalent metal ions resulted in immediate gelation with the formation of hydrogel beads when the solution droplets hit the gelation bath. When converted into aerogels, the beads showed unimodal size distributions with an average particle mean diameter of 1.41 ± 0.06 mm and sphericities in the range of SPH = 0.82 – 0.91 (**Fig. S2 – S3**). In contrast to gelation via bivalent ions, the use of Pt(IV) as a crosslinker led to formation of weak wet gels of an undefined shape during gelation and final dry aerogels of foam-like character (**Fig. S3 e)**). These observations are in line with previous works, which showed that the crosslinking of alginate with metals strongly depends on the hard-soft character of the ions.[24] The low capability of Pt(IV) to crosslink alginate chains may be attributed to the softer character of the ion as compared to the other crosslinking metal ions in this work, which are well coordinated by hard oxygen atoms in negatively charged carboxylate groups and –OH-groups of β-D-mannurate (M) and of α-L-guluronate (G) building blocks. Additionally, Pt(II) as well as Pt(IV) tend to form characteristically inert “inner-orbital” type complexes in general[25]: therefore, the reaction rates of complex formation change significantly with the type of the original ligand, as well as with additional factors, such as pH. In our case, we can conclude that either 1) the coordination of Pt-ions in the alginate matrix was not strong enough or 2) the ligand exchange with hydrated $PtCl_4$ was not fast enough to fix the shape of the alginate solution during gelation before it is dispersed. These effects are also indicated by the results of the ICP-OES analysis, which yielded

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by far the smallest metal mass with respect to the alginate mass (r_{metal}) in the Pt-containing aerogel (**Table 1**). The low r_{metal} value for Pt(IV) may result from either low affinity of Pt(IV) to the alginate matrix or washing out effects during the washing/solvent exchange steps as a result of weak coordination. Coordination of other metal ions resulted in a significant increase in r_{metal}. Aerogel skeletal density correlates with the ratio of r_{metal} and parent metal density (**Fig. S4, left**) (with exception of Ca-CA, since calcium-chloride and alginate show relatively similar parent densities). While no systematic relation between r_{metal} and textural properties was determined, the type of crosslinking ion has a strong effect on the overall mesopore volumes and specific surface areas. Gelation via comparably hard ions Ca(II), Ni(II) and Cu(II) resulted in higher specific surface areas as well as mesopore volumes as compared to Pd(II) and Pt(IV) crosslinked gel, whereas the trend of the specific surface area is in line with the affinity of different bivalent ions with alginate molecules.[26] Hereby, values of S_V (480 – 687 m^2/g) are in the typical range as reported for Ca(II), Cu(II) and Zn(II) crosslinked alginate and alginate-polyurea hybrid aerogels.[8,20,24, 27,28]

Table 1 Properties of M-CAs crosslinked with different metal ions at constant c_m = 17 mmol/L.

Material [-]	r_{metal} (wt%)	S_V (m^2/g)	d_{meso} (nm)	V_{meso} (cm^3/g)	ρ_s (g/cm^3)	ρ_b (mg/cm^3)	Total porosity (%)
Ca-CA	8.9 ± 0.5	494 ± 25	41.6	7.0 ± 0.4	1.71 ± 0.09	90.5 ± 0.3	94.7
Ni-CA	9.5 ± 0.2	480 ± 24	41.4	6.5 ± 0.3	1.9 ± 0.2	35.8 ± 0.2	98.1
Cu-CA	11.1 ± 0.2	687 ± 34	17.4	6.7 ± 0.3	2.3 ± 0.1	106.8 ± 0.1	95.1
Pd-CA	12.9 ± 0.6	243 ± 12	17.0	2.3 ± 0.1	4.6 ± 0.9	85.68 ± 0.04	98.1
Pt-CA	3.40 ± 0.04	268 ± 13	27.1	1.29 ± 0.01	1.69 ± 0.07	-	-

To the best of our knowledge, this work is the first example of purely Pt(IV) crosslinked alginate aerogels. We also determined a considerably higher specific surface area for Pd-CA (242 m^2/g) than reported elsewhere (52 m^2/g).[20] The low surface areas in the compared work could result from the presence of excess free metal salt in the alginate matrix (as also indicated by the high Pd content used of r_{metal} 17.6 wt%), which would lead to a decrease in S_V given the large difference in atomic

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mass of the metal compounds as compared to carbon. In this work, the decrease in S_V was also obtained, when increasing the metal salt concentrations of Ca(II), Ni(II) and Cu(II)-salts in the gelation bath above c_m = 17 mmol L^{-1} (**Fig. S4, right**) as well as in the work of Miller *et al.*, who obtained a decreased S_V with increasing metal loading in the Ru/carbogel nanocomposite.[29] In contrast to the trend of the specific surface area, no significant differences were found in the pore size distributions and mean pore diameters if c_m is increased above 17 mmol/L (**Fig. S5**). This shows that additional ions in the gelation bath do not additionally contribute to crosslinking and cause no microstructural changes.

Since no metal salt crystals or deposits were detected in SEM analysis in any substrates (**Fig. S6 - S10**) produced at c_m = 17 mmol/L, we conclude that metal ions were quantitatively integrated into the alginate matrix at the given concentration. Typical macro- to mesoporous (d_{pore} = 5 – 50 nm) fibrous microstructures are visible in the SEM pictures, with relatively broad meso pore size distributions and mean mesopore diameters d_{meso} being in the range of 17 – 42 nm (**Fig. 1**). Optical evaluation of SEM pictures shows, that crosslinking with Pd(II) and Pt(IV) yields comparatively loose networks with a significant fraction of macropores, whereas the Cu(II) crosslinked gel shows the densest network with less visible pores in the macroporous range (**Fig. 1 a) - e), Fig. S6 - S8**). The qualitative results from SEM are in line with results from gas sorption analysis, which yielded lower mesopore volumes for Pd(II)-CA and Pt(IV)-CA.

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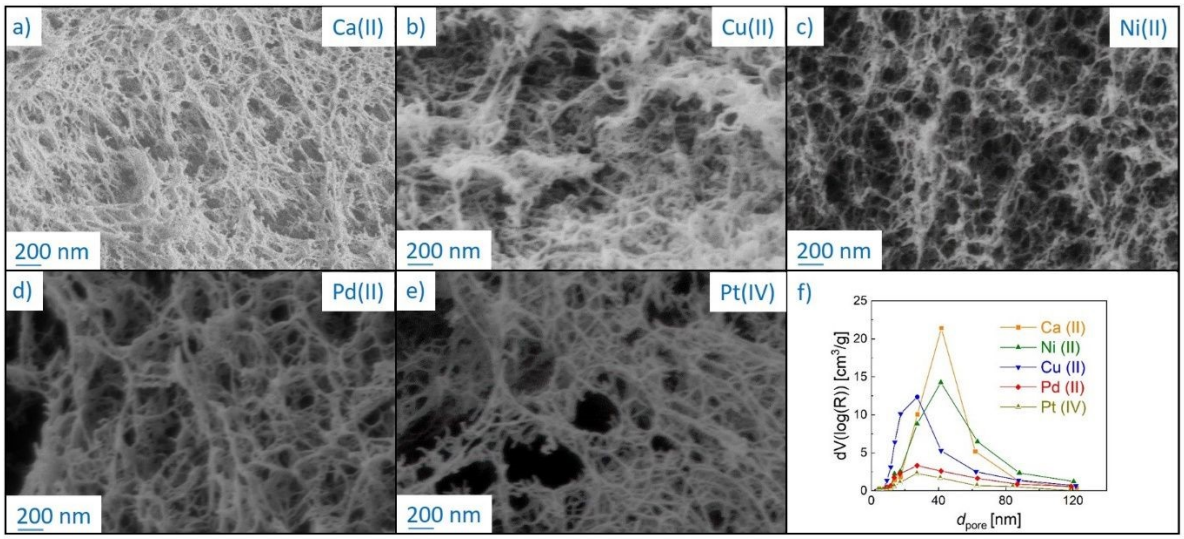

Fig. 1 a) – e) SEM pictures of inner pore structures of M-CAs crosslinked with different metal ions at c_m = 17 mmol/L and f) their pore size distributions. Lines are drawn to guide the eye.

Furthermore, a thin denser but still porous outer layer is identified in SEM pictures of aerogels beads (**Fig. S9 - S10**). This result is in line with previous works, and it is caused by the gelation mechanism. When sol droplets hit the gelation bath, the immediate complexation of alginate by metal ions on the surface of droplets leads to formation of a comparably dense, porous skin, followed by a slower, diffusion controlled completion of the reaction in the inner droplet part.[8] In summary, the microstructure of M-CAs is determined by the metal ion chosen for the crosslinking of alginate, while keeping the total porosity at a constant level. In the next step, the pyrolysis of the produced M-CAs was investigated for gels produced at c_m = 17 mmol/L.

Influence of pyrolytic treatment on mass loss and textural properties

Thermogravimetric analysis (**Fig. 2**, **a)**), and derivative thermogravimetric (DTG, **Fig. 2**, **b-c)**) curves were recorded to investigate the mass loss of M-CAs during pyrolysis in the whole temperature range up to final pyrolysis temperatures of T_P = 600 °C at a constant heat rate and to identify T_P values, at which the most significant changes occur. Signals in the DTG curves represent major mass loss and are used to determine the thermal behavior steps of the materials[30], whereas the individual peaks and maxima are defined in order of appearance as peak X and TX_{max} (with X = 1 – 3). The

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decomposition patterns of all samples can generally be divided into three distinct main steps: 1) moisture loss; 2) biopolymer degradation and 3) slow sustained degradation.[31]

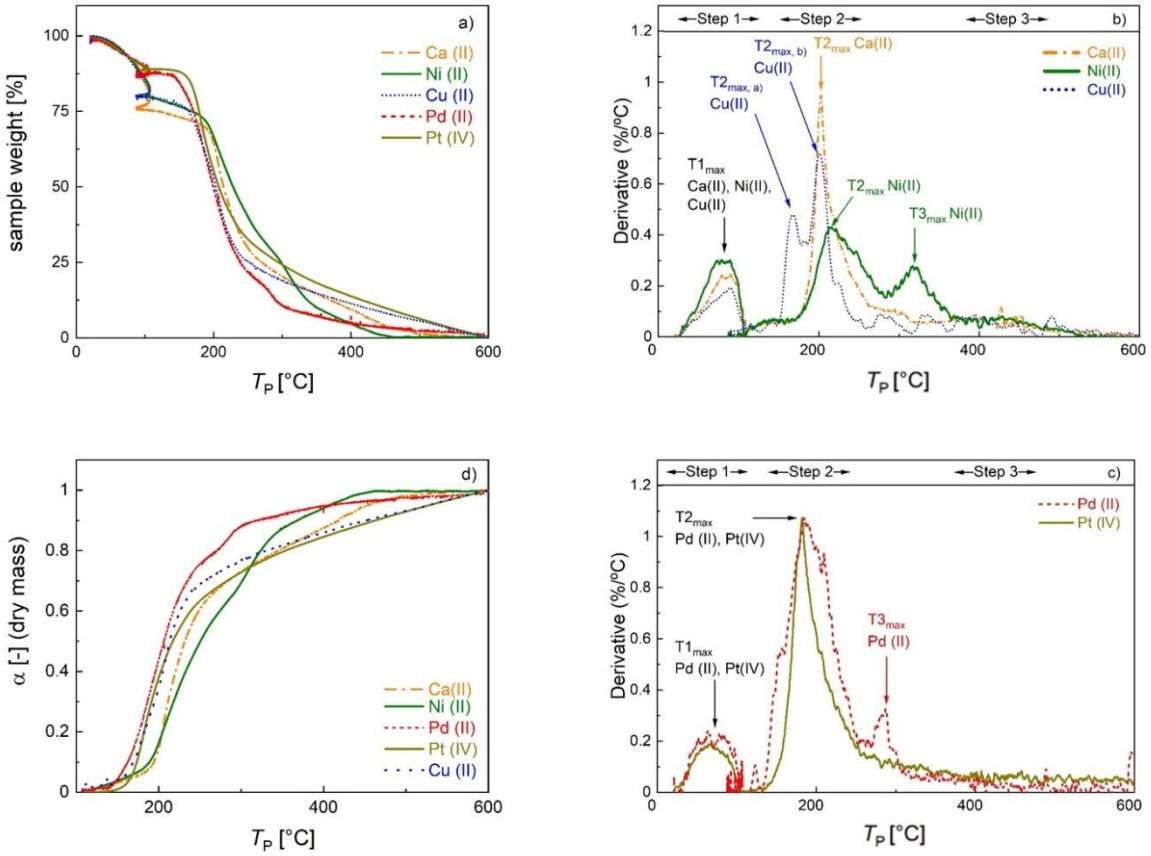

Fig. 2 Normalized TGA curves a), DTG curves b), c) and degree of conversion profiles d) of M-CAs crosslinked with different metal ions. See text for details.

Alginate aerogels crosslinked with different metal ions show a metal-specific decomposition profile (**Table 2**). The first DTG peak (T_P = 22 - 108 °C) is due to moisture loss. A relatively high mass loss (6 - 18 % with respect to the initial mass of the sample, denoted as Δm_{water} in **Table 2**) is rather expected at this stage given the hydrophilic and porous nature of the materials. The second DTG peak occurs in the range 130 – 255 °C for Ca-, Cu- and Pt-CAs and 130 – 364 °C for Ni- and Pd-CAs. The second peak corresponds to the major mass loss of 39 - 61% (denoted as Δm_d and calculated with respect to the dry mass, i.e. the mass after the first step). The onset temperature T_{onset} of peak 2 is assumed to mark the beginning of the thermal scission of the carboxylate/carboxylic acid

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groups via decarboxylation/decarbonylation and was determined to follow the order: Pd(II) ≈ Pt(IV) < Cu(II) < Ca(II) ≈ Ni(II). In general, the decomposition mechanism of metal-crosslinked alginates is complex and far from being completely understood. For instance, over 20 different degradation products were determined in the decomposition of Ni- and Cu-crosslinked alginate in each case, suggesting metal-specific decomposition mechanisms.[32,33]

Table 2 Thermal parameters in the pyrolysis of different M-CAs.

Material [-]	DTG-peak 1 [°C]	DTG-peak 2 [°C]	DTG-peak 3 [°C]	$T1_{max}$ [°C]	$T2_{max}$ [°C]	$T3_{max}$ [°C]	Δm_{water} peak1 [%]	Δm_d peak2 [%]	Δm_d peak3 [%]	Δm_d total [%]
Ca-CA	27 - 108	174 - 255	-	65	185	-	17.7	39.2	-	67.3
Ni-CA	22 - 108	174 - 283	283 - 365	79	215	318	17.1	46.0	21.8	81.3
Cu-CA	22 - 108	150 - 255	-	87	162[a)] 202[b)]	-	13.1	44.2	-	67.4
Pd-CA	22 - 107	131 - 255	267 - 309	68	186	285	12.2	41.6	5.7	54.5
Pt-CA	22 - 107	135 - 255	-	65	182	-	10.2	60.9	-	95.3

a), b): see assignment in Fig. 2.

The decomposition profiles observed for Ca(II),- Cu(II)-, and Ni(II) are consistent with previous works, whereas the T_{onset} of Ni(II)-crosslinked alginate is shifted to higher temperatures.[8,33] Differences in the values of T_{onset} and $T2_{max}$ compared to the cited works may be attributed to the alginate composition (G/M ratio, molecular weight) as well as to the different degree of crosslinking.[33] For the Cu(II)-crosslinked sample, a DTG duplet peak is recorded at 165 and 203 °C, which was also observed in other works.[8,31] In the decomposition of pure metal salts, the duplet peak at lower temperatures (45 - 58 °C) corresponds to the formation of trihydrate and monohydrate species of the pentahydrate sulfate.[34,35] However, in this work this phenomenon took place at higher temperature, suggesting a strong interaction of the Cu-complex with alginate chains. In contrast to strong coordinating ions, T_{onset} for alginate crosslinked with soft Pd(II) and Pt(IV) ions is shifted to lower temperatures. Furthermore, a third peak appeared at T_P > 265 °C for Ni(II)- and Pd(II)-crosslinked samples may be ascribed to the decomposition of the remaining metal precursor or to the further decomposition of alginate. In the

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case of Ni-CA, a relatively high mass loss (Δm_d > 20%) in this region points out rather biopolymer decomposition.[36,37]

The overall mass loss with respect to the dry mass ("Δm_d, total" in **Table 2**) covers a wide range (55–95 %). Literature results for ionic crosslinked alginate aerogels are close to the upper values of this range.[8,38] In our case, the variation in the total mass loss is linearly related to r_{metal} (**Fig. S11**), showing that the variations in overall yield can be attributed to the amount of non-pyrolyzed inorganic material. Due to significant differences in the total mass loss caused by variation of inorganic/organic component ratios, additional insight into effect of individual metal ions on biopolymer decomposition is provided by plotting the degree of conversion α against T_P (**Fig. 2**, **d)**), which is defined as:

$$\alpha = \frac{m_{0,d} - m_n}{m_{0,d} - m_{end}}$$

with m_0 is the initial dry mass after completion of the drying step (T_P = 107 – 108 °C), m_n is the mass at the corresponding temperature $T_{P,n}$, and m_{end} is the final sample mass at 600 °C.[31]

Because the degree of conversion α is normalized with respect to the final mass, samples with significantly different metal/metal oxide content (such as Pd-CA and Pt-CA) can be directly compared. It is notable that in the temperature range T_P < 310 °C the degree of conversion of Ni-CA is the lowest curve compared to other M-CAs, while no further decomposition of Ni-CA took place at T_P > 460 °C. Furthermore, the shape of the curves for Ca-, Cu and Pt-CAs is similar for high conversions (α > 63 %) which are readily achieved in the low temperature range (T_P < 270 °C), followed by a slower decomposition at higher conversions. Pd-CA shows both, a low T_{onset} as well as rapid decomposition in the low temperature range.

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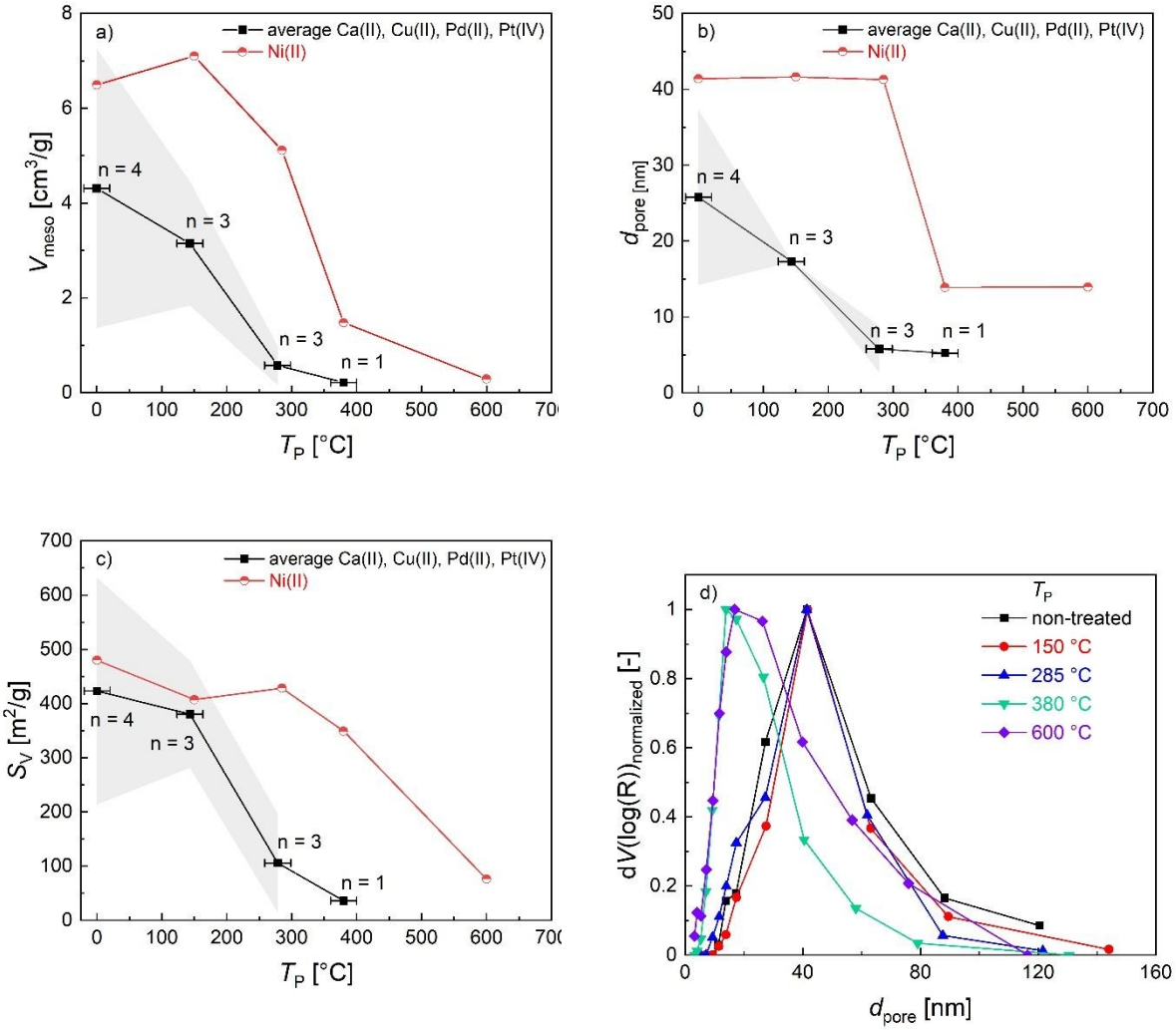

Fig. 3 Change of a) V_{meso}, b) d_{pore} and c) S_V in dependence of T_P. Squares represent the average values from aerogels crosslinked with different metal ions, excluding Ni(II). Lines are drawn to guide the eye, error areas represent the standard deviation from averaged values (*y*-error), n represents the number of single values taken into account varying at different temperatures, error bars represent the range of averaged values of T_P (*x*-error). Values at T_P = 0 °C represent non-pyrolysed M-CAs. d) Normalized pore size distributions of Ni-crosslinked gels heated up to different temperatures. Lines are drawn to guide the eye.

Since mass loss in aerogels pyrolysis is connected to textural properties, changes in the mesoporous structure during pyrolysis were followed via N_2 sorption analysis of samples, which were pyrolyzed up to four different T_P values whereas the pyrolysis process was stopped after the most significant mass loss steps respectively (T_P = 130 - 150 °C, 265 - 285 °C, 380 °C, 600 °C). The general trend is that the S_V and V_{meso} values decrease over the course of the pyrolysis process (**Fig. 3, a) – c)**), with

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the most significant changes being directly related to different steps of the mass loss. With the exception of Ni(II)-crosslinked aerogels, pyrolysis up to T_P = 285 °C results in a significant loss of mesoporosity. This is also reflected in the applicability of BET/BJH-methods: the M-CAs and materials obtained at low pyrolysis temperatures show typical type IV-isotherms and highly correlated BET-plots (**Fig. S12**, R^2 0.9991—0.9999), whereas determination of the textural properties via N_2-porosimetry is not possible for most M-DCs obtained at values of T_P > 300 °C. To date, the exact mechanisms behind the transformation of the intrinsic biopolymer-aerogel pore structure to the corresponding carbon network have not been revealed: due to the complexity of the pyrolysis process, it seems reasonable to assume that different effects might dominate the transition depending on the starting material and its intrinsic pore structure. From a fundamental point of view, the following processes might lead to the diminishing, shrinkage, or formation of porous voids in the system, e.g., a) partial collapse of the already existing macro- and mesoporous network resulting from mass loss in the form of gases. During this process, effects like pore-opening and pore-enclosure might occur in different temperature ranges of the pyrolytic treatment.[38] b) Creation of new pores by evaporation of material from the bulk phase. In case of the latter, the existence of a pore network in the starting material is not essential to obtain carbon aerogels, and pores could be formed by starting from a multi-component system followed by pyrolysis or subsequent etching of only one component (template-approach), as has e.g. been shown for polymer crosslinked silica-based xerogel powders.[39] In our case (one component system, if metal-ions are neglected), the mass loss during pyrolysis may result in transformation of the already existing pores. This is reflected in the mesoporous range by BJH-pore size distributions of Ni(II)-crosslinked gels: while no changes are observed up to T_P = 285 °C, further increase to T_P = 380 °C results in a shift in the mesoporesize-distribution to smaller pore sizes, being indicative for shrinkage of intrinsic mesopores (**Fig. 3, d**). It is interesting to note that this effect coincides with the $T3_{max}$-signal in the corresponding DTG-curve (**Fig. 2**, **b)**, showing that prior mass-loss of up to 46% ($T2_{max}$-signal) did not lead to substantial changes of the pore network. The additional increase in the pyrolysis temperature up to T_P = 600 °C then leads only to small changes in the textural properties (a slight broadening of the pore size

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distribution and an increase in the pore volume in the range of larger mesopores), which is consistent with the minor mass loss occurring in the temperature range between T_P = 380 – 600 °C (**Fig. 2**, **a, b)**. In summary, our results suggest that a slower and uniform conversion rate of organic material, as obtained for Ni-CA, is beneficial for the (partial) preservation of mesoporosity.

Influence of pyrolytic treatment on backbone morphology and evolution of nanoparticles

Pyrolysis of M-CAs resulted in pronounced volumetric shrinkage of beads (approx. 31 - 52%), whereas structural integrity was generally maintained (**Fig. S13**). However, in the case of Cu-DC and Pd-DC, some morphological changes are visible in SEM pictures, which indicate fast shrinkage or outgassing processes due to the comparably fast conversion in the lower temperature range (see **Fig. 2d**). In contrast, no defects in Ca-DC and Ni-DC beads were identified. SEM-BSE data provide a qualitative assessment of the morphological changes and especially the evolution of emerging nanoparticles during pyrolysis: as the intensity of BSE strongly depends on the atomic number, metal-containing structures are visible as brighter zones embedded in the darker biopolymer/carbon backbone (**Fig. 4**).

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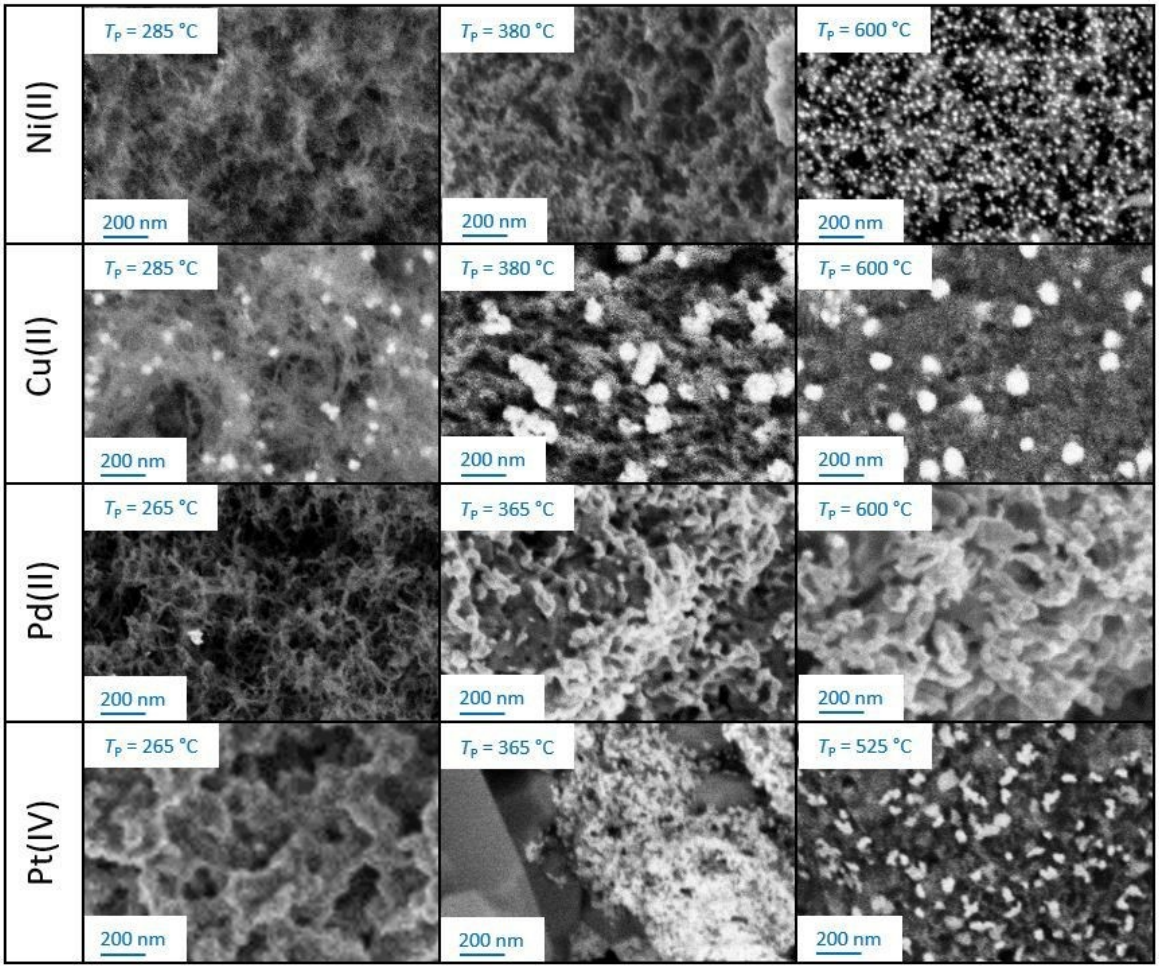

Fig. 4 Exemplary SEM-BSE images of M-DCs obtained after stopping the pyrolysis temperature at three different temperatures T_P.

Besides the evolution of the nanoparticles, changes of the backbone can also be observed: all materials obtained at T_P = 285 °C still show a fibrous structure that is typical for biopolymer aerogels (Pt-crosslinked aerogel represents an exception). At higher pyrolysis temperatures, the backbone demonstrates further densification and reduction in pore size. While all aerogels remain porous upon pyrolysis, the pore densification process during pyrolysis is highly influenced by the crosslinking metal. Optical evaluation of SEM-BSE pictures allows for a rough categorization: the best resistance against temperature-induced morphological changes is observed in the following order: Ni(II) > Cu(II) > Pd(II) > Pt(IV). This qualitative result highlights the crucial influence of crosslinking metals ion on the materials resulting from carbonization process.[8] The result is also in line with BET/BJH and TGA

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finding as well as with former work on Ni- and Cu crosslinked alginate.[33] In the following, trends in the evolution of nanoparticles are described for each crosslinking ion. An overview of all results is provided in **Table 3**.

Table 3 Average particle size/species of metal NPs in different M-DCs determined via SEM-BSE and XRD.

M-DC	T_P (°C)	Species [-]	d_{NP} estimated via image analysis [nm]	Crystallite size estimated via XRD [nm]
Ni-DC	380	NiO	–	8 ± 1
Ni-DC	600	NiO	14 ± 7	23 ± 2
Cu-DC	285	Cu	40 ± 19	15 ± 5
Cu-DC		Cu_2O		23 ± 15
Cu-DC	600	Cu	85 ± 29	21
Cu-DC		CuO		25 ± 6
Pt-DC	525	Pt	33 ± 23	25 ± 5

Ni-carbogels

Below T_P = 380 °C no NPs are observed neither in BSE-SEM nor in XRD. After pyrolyzing the aerogels up to T_P = 380 °C, the BSE image shows tiny, well-defined NPs (**Fig. 4**); however, determination of the NP size is not possible due to low contrast. Bragg's peaks of nickel(II)-oxide (NiO) were detected by XRD (**Fig. 5, a**) with an average crystallite size of approximately 8 ± 1 nm. Pyrolysis at T_P = 600 °C also results in the formation of NiO-NPs. The average NP size is determined to be 14 ± 7 nm from the BSE images (**Fig. 4**), whereas significantly larger crystallite sizes of NiO NPs of (23 ± 2 nm) were detected based on narrow NiO Bragg peaks (**Fig. 5, a**). The diffraction peaks at 44.52, 51.83 and 76.34° are indicative of the face-centered cubic (fcc) phase of crystalline Ni (JCPDS no. 04–0850) corresponding to (111), (200), and (220) planes of crystalline elementary Ni.[41] At a pyrolysis temperature of T_P = 600 ºC, these peaks become more detached from the NiO peak. NiO peaks are still present but weaker than at T_P = 380 ºC. This decrease in intensity can be attributed to the growth of crystallite size in both NiO and Ni, resulting in narrower peaks in the XRD pattern. Rietveld analysis was used to estimate the composition of Ni and NiO. It was shown that the amount

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of Ni significantly decreased during high-temperature pyrolysis, i.e. the ratio of NiO and Ni was 90.4% : 9.6% at 380 °C vs. 95.4% Ni : 4.6% at 600 °C. The temperature increase can lead to the sintering of metallic nickel into larger crystallites of metal, driven by the transformation of different crystal forms (e.g., 111, 200, 220) of nickel..[42]

Cu-carbogels

Cu containing aerogels pyrolyzed at different temperatures lead to the formation of NPs of different chemical compositions at lower pyrolysis temperatures than for Ni-CA. Bragg's peaks of elementary copper and copper (I) oxide were observed in samples pyrolyzed up to T_P = 285 °C, however, copper (II) oxide dominates (with only traces of elementary copper detected) in the NPs after pyrolysis up to T_P = 600 °C (**Fig. 5, b**). Pyrolysis of Cu-CA under comparable conditions resulted in metallic Cu and Cu_2O crystallites, as a result of the redox reaction between Cu(II) and sugars, suggesting that during pyrolysis a carbothermal reduction of the metal ion occurred.[8] In the work by Baumann *et al*., metal particles with no sign of metal-oxides in the XRD patterns were formed upon pyrolysis of resorcinol-formaldehyde aerogels containing M(II) cations (Ni, Co, Cu).[11] The reason for the complete reduction in this case, and only partial reduction of the present work, can be that the resorcinol-formaldehyde resin is a reductive matrix, while the alginate is not necessarily. Besides oxidative state, the size of Cu-NPs also changed over the course of the pyrolysis: analysis of BSE images obtained at T_P = 285 °C shows that the particle size of Cu/Cu_2O NPs is 40±20 nm (**Fig. 4**), while the size of CuO/Cu NPs obtained at T_P = 380 and 600 °C is significantly larger as compared to NiO NPs: 85 ± 29 nm. The average crystallite size for each component of the NPs was calculated from the XRD data. A significant discrepancy can be seen between XRD and BSE data for the particle size. The reason is that the XRD can estimate the crystallite size of the NPs and not the physical size of the NPs. The general trend that the crystallite size increases with increasing pyrolyzing temperature holds true for all components (see **Table 3**). According to the result of the Rietveld analysis, we found the following: the ratio of elementary Cu and Cu_2O is 51.6% and 48.4% in the case of low-temperature

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pyrolysis, however, a significantly lower amount of Cu(0) has been detected after high-temperature pyrolysis, with a 0.7% Cu to 99.3 % CuO ratio.

Pd-carbogels

In the case of Pd-DC, no formation of defined nanoparticles was observed via SEM-BSE. Brighter zones of samples produced at $T_P \geq 365$ °C indicate the formation of Pd-layers around the carbon backbone, most probably due to the high final metal content (**Fig. 4**). XRD data show that the metal part of the sample is in a crystalline state after pyrolyzing at $T_P = 265$ °C, i.e. that 63.4% Pd and 36.6% PdO can be estimated by the Rietveld analysis (**Fig. 5, c**). This ratio changed slightly for the samples pyrolyzed at $T_P = 600$ °C, resulting in a ratio of 46.9% Pd and 53.1% PdO.

Pt-carbogels

NPs could be detected in the carbogel only after pyrolyzing at 525 °C. However, the shape of the NPs is not well-defined compared to those of the Ni- and Cu-carbogels. The BSE image of the Pt-carbogel sample clearly shows (**Fig. 4**) the appearance of large, irregular shaped NPs. XRD measurements also confirmed that elementary Pt is the dominant component of the NPs (**Fig. 5, d**). Additionally, Bragg's peaks of sodium chloride were also detected, which was expected due to its low solubility in ethanol in the solvent exchange step and therefore reaction of Na-ions from sodium alginate and chlorine from $PtCl_4$ during pyrolysis. The size of Pt-NPs is 33 ± 23 nm according to the image analysis of BSE pictures and in agreement with the evaluation of the XRD data (**Table 3**).

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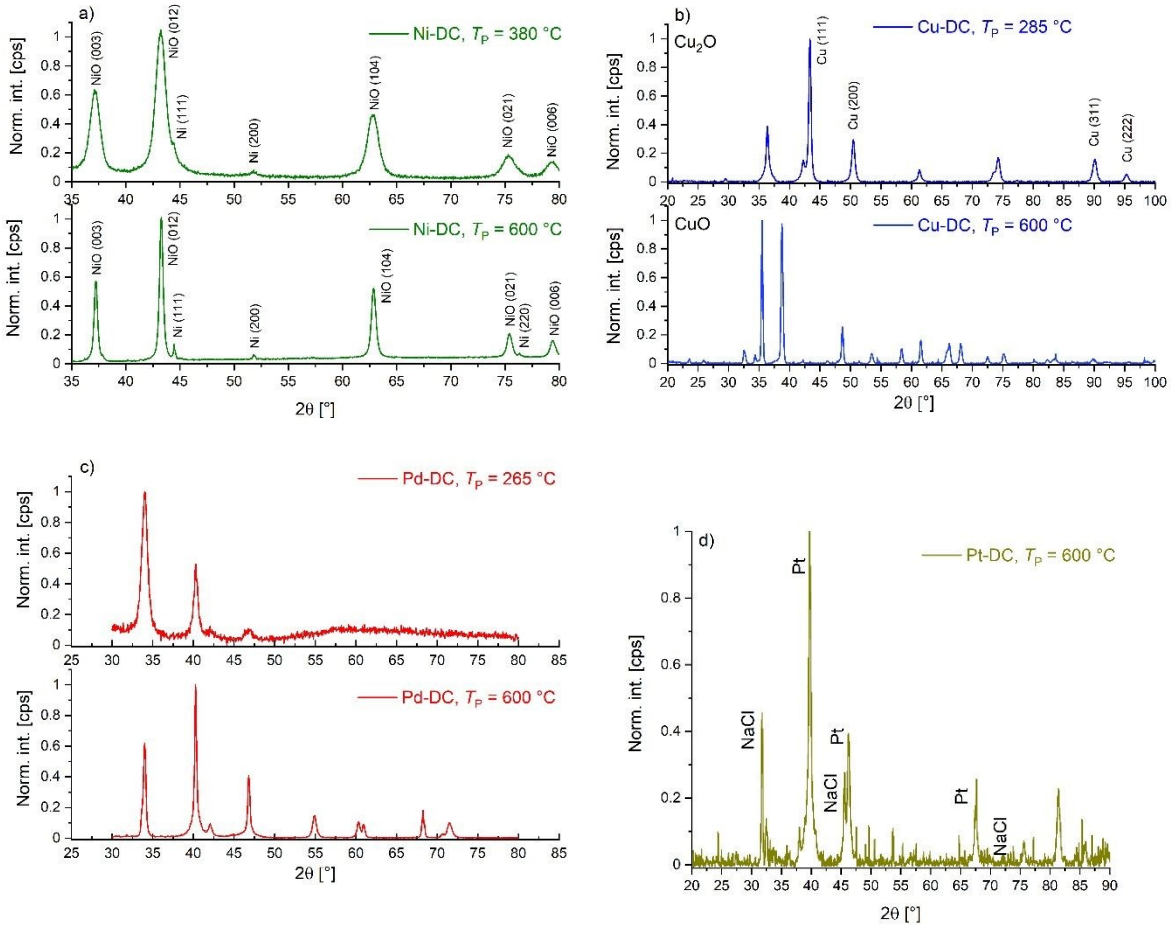

Fig. 5 XRD patterns of different metallic metal/metal oxide nanoparticles obtained from different M-Cas and different values of T_P: a) Ni-DC, b) Cu-DC, c) Pd-DC, d) Pt-DC.

Reducibility of metal oxides

The temperature programmed reduction (TPR) technique allows determining the reduction temperature of metal oxides, revealing its reducibility, which is an important parameter to determine the optimal activation conditions for potential applications.[43] H_2-TPR of Ni-DC (T_P = 600 °C) (**Fig. 6**) showed a main broad peak at 334.3 ºC, characteristic of the reduction region of Ni(II) to Ni (0) (378 ºC for bulk NiO, and 382 ºC for a Ni4.3%/AC catalyst), confirming that there is only one oxidative state in Ni-DC.[44,45] The slightly lower reduction temperature (334.3 ºC) may be related to the higher metal content in the Ni-DC (9.5%), also reflected by the broad shape of the reduction peak, thus providing a weak surface level interaction between Ni and hydrogen. This higher reducibility is comparatively competitive to other Ni based catalysts with similar metal content and lower reducibility, such as

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Ni10%/AC (main peak at 430 ºC).[46] This imply an easier activation pretreatment of Ni-DC would be needed in further applications. Moreover, the Ni-DC did not present significant gasification of the carbon lattice, that usually onsets from 600 ºC for carbon supports, hence supporting evidence of its thermal stability.[47,48]

Fig. 6 also illustrates the reduction of copper at different oxidative states, observing two major steps. The first one starts developing from 60 ºC and is completely completed at 420 ºC, corresponding to the reduction of CuO species (Cu(II)) to Cu_2O (Cu(I)) and Cu(0).[46] The second one onsets at 520 ºC, with a maximum at 673 ºC. Since copper oxide species completely reduce to metallic Cu(0) at temperatures no further than 420 ºC, this second step should be explained by the gasification of the carbon backbone of the Cu-DC sample.[49,50]

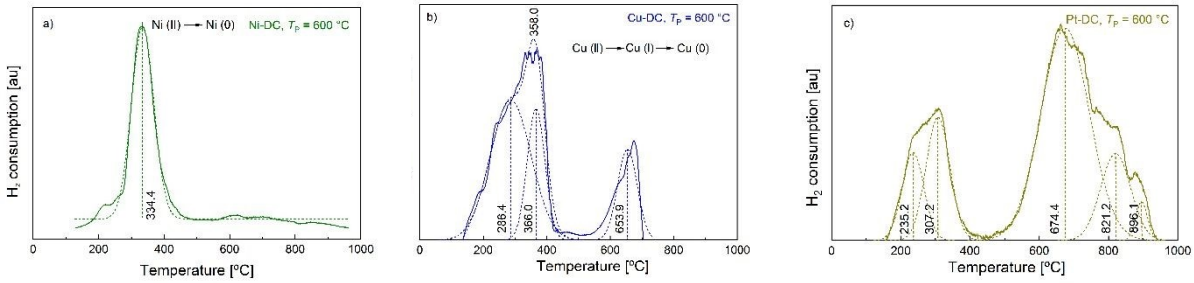

Fig. 6 H_2-TPR curves of M-DCs produced at T_P = 600 °C**.** Straight lines represent raw data, dashed lines the deconvoluted peaks and according maxima.

The deconvoluted reduction peaks of Pt-DC at 235.2 and 307.2 ºC (**Fig. 6 c)**) can be attributed to the reduction of bulk and inter surficial PtO_xCl_y, as suggested by Hwang et al., originating from traces of the $PtCl_4$ precursor.[51] Furthermore, HCl desorption has also been identified within this temperature range.[52] It appears that there is an absence of typical PtO_x reduction peaks (expected to be below 200 ºC),[53] probably associated with the poor crosslinking that displayed the Pt-CA. This means that unlike Ni, Cu and Pd, Pt does not catalyze the oxide formation during carbonization. Finally, gasification of the Pt-DC sample is observed from 450 to 950 ºC. The TPR curves of Pd-DC where not possible to obtain with the method used, in order to avoid serious damage on the chemisorption equipment, due to the detection of condensable substances. In summary, H_2-TPR analysis shows

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that among the tested M-DCs, Ni-DC has the best performance toward activation by the H_2 reduction process, because it presents: i) only one single reduction step, ii) the lowest reduction temperature, iii) the highest thermal stability of the carbon backbone as indicated by a minimal gasification.

CONCLUSIONS

This work presents a facile method for the production of nanoparticle-decorated porous carbons by pyrolysis of aerogels obtained via crosslinking of sodium alginate with different metal ions and supercritical drying. The broad variety of textural properties obtained for parent aerogels emphasizes, that choice of crosslinking metal ions is a suitable tool for manipulating the pore structure in alginate aerogel synthesis. In the pyrolysis step up to 600 ºC, all M-DCs showed a significant mass loss and loss a of mesoporosity, whereas the highest stability was identified for the Ni-crosslinked gel with the lowest conversion rate during pyrolysis, which shows the significant influence of the metal ion used on the pyrolysis mechanism. XRD and BSE measurements show the growth of metal(oxide) nanoparticles from the intrinsically molecularly distributed metal ions and the possibility of generating a carbon matrix with homogeneously distributed NPs. While the advantages of this approach lie clearly in its green nature and the avoidance of costly and time consuming post-processing steps (like e.g. wet-impregnation our supercritical deposition) in order to distribute NPs in the carbon matrix, it has to be pointed out, that limitations are given by the fact, that the properties of nanoparticles and porous matrix cannot be changed independently. In order to achieve desired nanoparticle sizes with metals of choice, individual strategies have to be developed which could include the use of crosslinking metal ion mixtures, variation of alginate and metal-precursor concentrations and the use of alginates with different M- to G-ratios. Compared to the properties of conventional carbogels based on resorcinol-formaldehyde, the M-DCs developed in this work showed significantly lower specific surface area and higher mass loss during pyrolysis.[54] The optimization potential becomes evident when considering that the pyrolysis temperatures employed in this study remain below the threshold at which micropore formation might occur. Hence, the application of a higher $T_P \geq 800$ °C is a

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promising next step, which is expected to result in the formation of additional micropores. Enhanced preservation of the mesopore volume/specific surface area may be achieved by employing lower heating rates during pyrolysis (e.g., in the range of 1.0 – 2.5 °C/min), as has already been reported[39] and is also indicated by our results with the lowest conversion rate yielding comparably best textural features of the carbon matrix. A further possibility to reduce mass loss and to improve textural properties is the introduction of lignins or tannins into the alginate matrix as both biopolymers have reactivity and properties similar to those of resorcinol and phenol and have already successfully been used as non-toxic replacement with natural origin.[55] We believe that studies conducted in this work are helpful in identifying suitable conditions and might therefore contribute to the development of catalyst-systems based on green synthesis routes.

Author contributions

Juan I. del Río: investigation, original draft, formal analysis. Laura Juhász: investigation, formal analysis. József Kalmár: conceptualization, formal analysis, editing. Zoltán Erdélyi: investigation. María D. Bermejo: investigation. Ángel Martín: investigation. Irina Smirnova: resources. Pavel Gurikov: project administration, formal analysis, editing, Baldur Schroeter: conceptualization, original draft, formal analysis.

Conflicts of interest

There are no conflicts to declare.

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